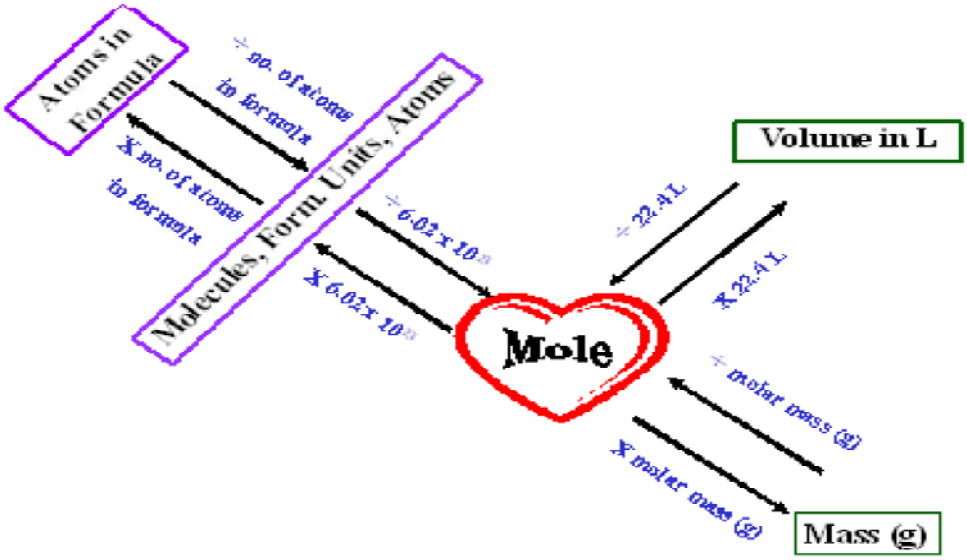
Name: _____ **Class:** _____ **Date:** _____ **ID: A**

Practice test chapter 10

Multiple Choice - *Choose the correct answer for the following questions. You may write on the test, only your answers on the scantron will be graded.*

**You will have questions over past chapters 1,2,3,4,25, 5, 6, 7, 8 and lots of naming from chapter 9.*



____ 1. How many moles of CH_4 are contained in 96.0 grams of CH_4?

a. 3.00
b. 12.0
c. 6.00
d. 16.0

____ 2. What type of representative particle is H_2O?

a. formula unit
b. molecule
c. atom

____ 3. Look at the compound below. How many total atoms are present for each element?

$Zn(ClO)_2$

a. 4
b. 5
c. 6
d. 3

____ 4. How many atoms are contained in 97.6 g of silver?

a. 5.45×10^{23}
b. 3.01×10^{23}
c. 1.20×10^{24}
d. 1.10×10^{23}

____ 5. Name the following Cl_2O_7

a. Perchlorate
b. dichlorine heptaoxide
c. dichlorine hexoxide
d. dichlorine heptoxide

____ 6. What is the gram formula mass of chromium (III) oxalate?

a. 65 g/mol
b. 192 g/mol
c. 396 g/mol
d. 368 g/mol

____ 7. Choose the correct formula for Ammonium oxalate.
a. $NH_4C_2O_4$
b. $(NH_4)_2C_2O_4$
c. $C_2O_4(NH_4)_2$
d. $(NH_4)_2C_2H_3O_2$

____ 8. What is the density of 1 mole of CO_2 gas at STP?
a. 1.34 g/L
b. 0.4 g/L
c. 1.96 g/L
d. 0.49 g/L
e. 0.75 g/L22

____ 9. What is the volume of 1 mole of Chlorine gas at standard temperature and pressure?
a. 22.4 L
b. 12.2 L
c. 6.02×10^{23} L
d. 67.2 L

____ 10. Choose the correct name for HF.
a. hydrogen fluoride acid
b. monohydrogen monofluoride
c. hydrofluoric acid
d. hydrogen monofluoride

____ 11. Calculate the number of molecules in 60.0 g of NO_2.
a. 7.85×10^{23}
b. 1.20×10^{24}
c. 7.9×10^{24}
d. 3.6×10^{25}

____ 12. Which compound/s represent a molecular compound?
a. CrO
b. CH_4
c. KNO_2
d. LiF

____ 13. How many oxygen atoms are there in 1.75 mol of Calcium phosphite?
Hint: First write the formula for calcium dichromate.
a. 3.01×10^{24} oxygen atoms
b. 6×10^{24} oxygen atoms
c. 7 oxygen atoms
d. 6×10^{25} oxygen atoms

____ 14. NaCl can be classified as what type of compound?
a. ionic
b. acidic
c. molecular

____ 15. SO_2 can be classified as what type of compound?
a. ionic
b. acidic
c. polyatomic ion
d. molecular

____ 16. How many atoms of chromium are there given 13 grams of chromium?
a. 1.5×10^{23}
b. 1.9×10^{26}
c. 3.3×10^{23}
d. 2.4×10^{24}

____ 17. Which of the following is a monatomic gas at STP?
a. Chlorine
b. Argon
c. Fluorine
d. Nitrogen

____ 18. What is the volume of a 200-gram sample of nitrogen gas at STP?
a. 320-liters
b. 250-liters
c. 160-liters
d. 125-liters
e. 8.9-liters

____ 19. How many representative particles are in 1.50×10^{23} moles of nitrogen dioxide?
a. 8.00×10^{-22} formula units.
b. 0.208 molecule
c. 1.25×10^{21} formula units.
d. 1.50×10^{23} atom
e. 9.03×10^{46} molecule

____ 20. What is the mass of a 55-Liter sample of krypton gas at STP?
a. 205-grams
b. 250-grams
c. 550-grams
d. 610-grams

____ 21. What is the correct formula for copper (I) peroxide?
a. CuO
b. CuO_2
c. Cu_2O
d. Cu_2O_2

____ 22. What is the percent composition of chromium in Barium chromate?
a. 4.87%
b. 9.47%
c. 20.5%
d. 25.2%

____ 23. What are the correct values for STP?
a. 1°C and 101.3 kPa
b. 0°C and 101.3 kPa
c. 0°C and 22.4 kPa
d. 0°C and 100 kPa

Multiple Response
Identify one or more choices that best complete the statement or answer the question.

____ 24. Which of the following is NOT classified correctly as a representative particle?
a. SO = molecular
b. Cu = formula unit
c. $(NH_4)_2SiO_3$= molecule
d. O_2= atom

Practice test chapter 10
Answer Section

MULTIPLE CHOICE

1. ANS: C PTS: 1
2. ANS: B PTS: 1
3. ANS: B PTS: 1
4. ANS: A PTS: 1
5. ANS: D PTS: 1
6. ANS: D PTS: 1
7. ANS: B PTS: 1
8. ANS: C PTS: 1
9. ANS: A PTS: 1
10. ANS: C PTS: 1
11. ANS: A PTS: 1
12. ANS: B PTS: 1
13. ANS: B PTS: 1
14. ANS: A PTS: 1
15. ANS: D PTS: 1
16. ANS: A PTS: 1
17. ANS: B PTS: 1
18. ANS: C PTS: 1
19. ANS: E PTS: 1
20. ANS: A PTS: 1
21. ANS: D
    ST 2A, 2B

    PTS: 1
22. ANS: C PTS: 1
23. ANS: B PTS: 1

MULTIPLE RESPONSE

24. ANS: B, C, D PTS: 1